

GCE

Chemistry B (Salters)

H033/01: Foundations of chemistry

Advanced Subsidiary GCE

Mark Scheme for November 2020

OCR (Oxford Cambridge and RSA) is a leading UK awarding body, providing a wide range of qualifications to meet the needs of candidates of all ages and abilities. OCR qualifications include AS/A Levels, Diplomas, GCSEs, Cambridge Nationals, Cambridge Technicals, Functional Skills, Key Skills, Entry Level qualifications, NVQs and vocational qualifications in areas such as IT, business, languages, teaching/training, administration and secretarial skills.

It is also responsible for developing new specifications to meet national requirements and the needs of students and teachers. OCR is a not-for-profit organisation; any surplus made is invested back into the establishment to help towards the development of qualifications and support, which keep pace with the changing needs of today's society.

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

© OCR 2020

Annotations

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

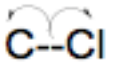
Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

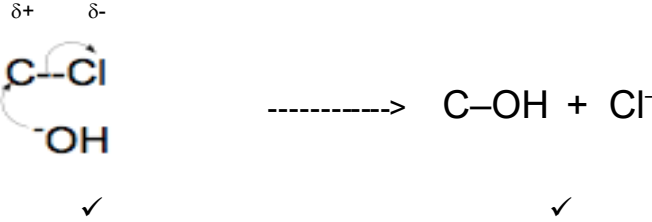
Annotation	Meaning
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

Section A

Question	Key	AO element
1	D	1.1
2	A	2.1
3	B	1.1
4	C	1.2
5	A	1.2
6	A	1.2
7	C	2.3
8	D	2.7
9	D	1.2
10	C	2.1
11	D	1.1
12	C	2.1
13	D	1.1
14	B	1.2
15	A	1.2
16	B	2.7
17	D	1.1
18	C	2.1
19	C	2.6
20	B	2.7

Section B

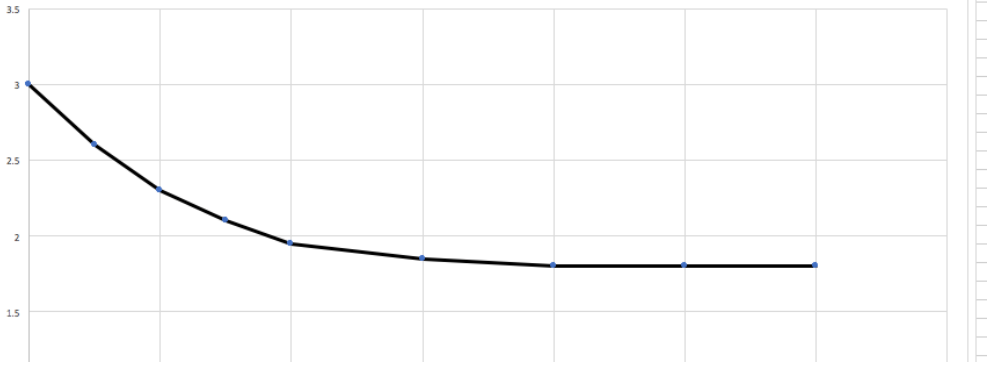
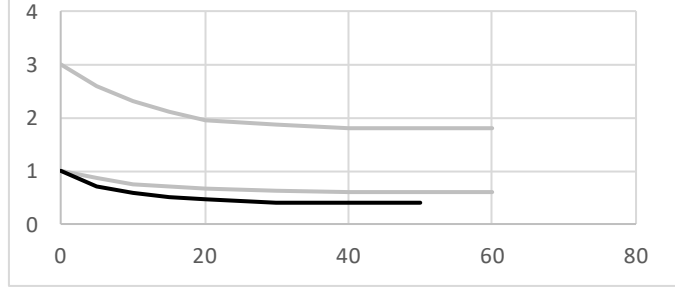
Question			Answer	Mark	AO	Guidance
21	(a)		dichlorodifluoromethane ✓	1	1.2	IGNORE spaces, and other separators
21	(b)	(i)	causes skin cancer/mutations OR damages crops ✓	1	1.1	ALLOW eye damage NOT eye problems
21	(b)	(ii)	It causes photochemical smog ✓	1	1.1	ALLOW toxic/poisonous/respiratory/breathing problems OR damage to plants/rubber
21	(b)	(iii)	Bonds vibrate (more) ✓	1	1.1	ALLOW They vibrate (more) NOT Atoms vibrate (more) IGNORE reference to collisions
21	(c)	(i)	$\text{ClO} + \text{O} \rightarrow \text{Cl} + \text{O}_2$ ✓	1	1.2	
	(c)	(ii)	$\text{Cl} + \text{Cl} \rightarrow \text{Cl}_2$ OR $2\text{Cl} \rightarrow \text{Cl}_2$ OR $2\text{ClO} \rightarrow \text{Cl}_2 + \text{O}_2$ ✓	1	2.1	
21	(c)	(iii)	Both 'propagation' ✓	1	1.1	
21	(d)	(i)	 AND homolytic (fission) ✓	1	1.2	NB Half arrows
21	(d)	(ii)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 3.46×10^{-5} (cm) award 4 marks Use of $v = E/h$ or implied by correct evaluation step(s) ✓ $v = 346000 / 6.63 \times 10^{-34} \times 6.02 \times 10^{23}$ (or correct evaluation 8.67×10^{14}) ✓ $\lambda = 3.00 \times 10^8 / 8.67 \times 10^{14} (= 3.48 \times 10^{-7} \text{ m})$ ✓ $= 3.46 \times 10^{-5} \text{ (cm)}$ ✓	4	2.2	ALLOW 2 or more sf. ALLOW ecf
	(e)	(i)	molecule/negatively charged ion with a (lone) pair of electrons which it donates(AW) to a (positively charged) atom (to form a covalent bond). ✓	1	1.1	

Question			Answer	Mark	AO	Guidance
	(e)	(ii)		2	2.5	One arrow (double headed) must start (when projected back) at bond and end (when projected forward) on Cl AND partial charges on C and Cl Other arrow (double headed) must start (when projected back) at minus (or a lone pair on OH) and end (when projected forward) on C IGNORE other atoms bonded to C. If both charges omitted can award second mark for balanced equation.

Question			Answer	Mark	AO	Guidance
22	(a)		6 protons; 7 neutrons ✓	1	1.2	
	(b)	(i)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 12.01 award 2 marks (98.9 x 12) + (1.1 x 13) ✓ divide by 100 and answer to 2 dp (12.01) ✓	2	2.2	ALLOW ECF
	(b)	(ii)	chance of 2 ¹³ C small (AW) ✓	1	3.2	
	(c)		 ✓ $M_r = 60$ (from M^+ peak in MS) ✓ C_3H_8O/C_3H_7OH ✓ CH_2OH only found in the primary isomer OR CH_2OH means OH at end (AW) ✓	4	3.1 3.2 3.1 3.2	ALLOW "60-31=29; which can only be CH_3CH_2 "

Question			Answer	Mark	AO	Guidance
23	(a)		(Otherwise) they react ✓	1	3.3	CON reactions with other substances
23	(b)	(i)	Na ⁺ AND its oxidation number goes down/goes from +1 to zero OR it gains electrons ✓	1	2.1	ALLOW 'sodium ion' NOT 'sodium'
23	(b)	(ii)	$2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$ OR $2\text{Cl}^- - 2\text{e}^- \rightarrow \text{Cl}_2$ ✓	1	2.4	ALLOW equation halved ALLOW 'e' without minus
23	(b)	(iii)	breathing apparatus (AW) ✓	1	1.1	ALLOW use in a fume cupboard ALLOW well ventilated room NOT face masks
23	(b)	(iv)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 1.4×10^5 (m³) award 5 marks (moles Na =) $1 \times 10^6/23$ (= 43478) ✓ moles Cl ₂ = half Na (21739) ✓ Rearrangement $V = nRT/P$ ✓ substitute values V (= $21739 \times 8.314 \times 873/1100$) = $1.43 \dots \times 10^5$ (m ³) ✓ 2sf and standard form ✓	5	2.8	ALLOW ecf Earlier points can be scored by implication in later ones, eg MP1 and MP2 from 21.74 in MP4; MP3 from correct expression in MP4 etc Award last MP for any number to two sf and standard form resulting from a shown calculation.
23	(c)		FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 1.1 or 1.07 award 2 marks 33/58.5 OR 0.56(4) (mole Na) AND 67/111(.1) OR 0.60(3)(mole Ca) ✓ ratio(= 0.60(3)/0.56(4)) = 1.1/1.07 ✓	2	2.6	ALLOW ecf
23	(d)		Na – sodium ions/(1)+ ions ✓	5	1.2	ALLOW labelled diagrams for all marks

Question			Answer	Mark	AO	Guidance
			delocalised electrons (AW)✓ NaCl – <u>Na^+ and Cl^-</u> ions ✓ 'lattice' or one structure point (eg 'alternating') ✓ Electrostatic forces (between oppositely charged ions) ✓			ALLOW opposite charges of ions attract
23	(e)	(i)	(colourless/pale green to) brown/orange/yellow ✓ $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$ ✓	2	2.5	ALLOW these colours alone or in any combination but no others. ALLOW ionic equation IGNORE state symbols
23	(e)	(ii)	EITHER add organic solvent – purple colour OR heat solution – purple vapour✓	1	3.4	

Question	Answer	Mark	AO	Guidance
24 (a)	Plotting points ✓ Smooth line of best fit ✓ 	2	2.8 3.3	ALLOW best fit line +/- half square from each point AND point 4 (2.20) off the line IGNORE labelling of line
24 (b)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = $0.046/4.6 \times 10^{-2}$ award 3 marks Expression for $K_c = [\text{CH}_4][\text{H}_2\text{O}]/[\text{CO}][\text{H}_2]^3$ ✓ Reading values from graph ✓ Calculation $(0.4)^2/0.6(1.8)^3 = 0.0457\dots$ ✓	3	2.8	ALLOW 2 or more sf. ALLOW ecf from expression or graph reading MP1 can be inferred from later steps IGNORE units
24 (c)	 line begins at or above 1 ✓ and flattens below 0.6 (but not to zero) ✓	2	3.1	Line can go below dotted line (or not)

Question			Answer	Mark	AO	Guidance
24	(d)		Heterogeneous AND catalyst and reactants in different states✓	1	1.1	

OCR (Oxford Cambridge and RSA Examinations)
The Triangle Building
Shaftesbury Road
Cambridge
CB2 8EA

OCR Customer Contact Centre

Education and Learning

Telephone: 01223 553998

Facsimile: 01223 552627

Email: general.qualifications@ocr.org.uk

www.ocr.org.uk

For staff training purposes and as part of our quality assurance programme your call may be recorded or monitored